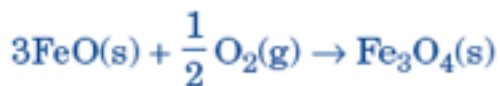
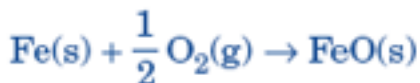


Hess' Law Worksheet

- 1 Calculate ΔH° for the reaction:



given the following data.



$$\Delta H^\circ = -272 \text{ kJ mol}^{-1}$$

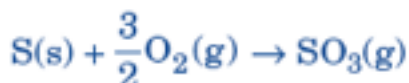


$$\Delta H^\circ = -1118 \text{ kJ mol}^{-1}$$

- 2 Calculate ΔH° for the reaction:



given the following data.



$$\Delta H^\circ = -395.2 \text{ kJ mol}^{-1}$$



$$\Delta H^\circ = -198.2 \text{ kJ mol}^{-1}$$

- 3 Calculate ΔH° for the reaction:



given the following data.



$$\Delta H^\circ = +67.7 \text{ kJ mol}^{-1}$$

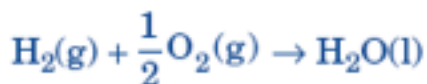


$$\Delta H^\circ = +9.7 \text{ kJ mol}^{-1}$$

- 4 Calculate ΔH° for the reaction:



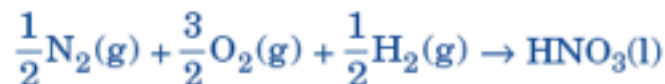
given the following data.



$$\Delta H^\circ = -285.8 \text{ kJ mol}^{-1}$$



$$\Delta H^\circ = -76.6 \text{ kJ mol}^{-1}$$

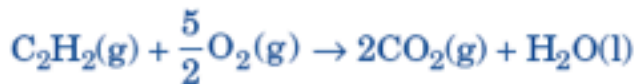


$$\Delta H^\circ = -174.1 \text{ kJ mol}^{-1}$$

- 5 Calculate ΔH° for the reaction:



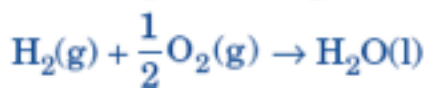
given the following data.



$$\Delta H^\circ = -1300 \text{ kJ mol}^{-1}$$



$$\Delta H^\circ = -394 \text{ kJ mol}^{-1}$$



$$\Delta H^\circ = -286 \text{ kJ mol}^{-1}$$