

# Structure of the Atom

Atoms are made up of three sub atomic particles: \_\_\_\_\_

Particle	Symbol	Location	Mass	Charge
Electron				
Proton				
Neutron				

To represent subatomic particles in an atom, use Standard Atomic Notation

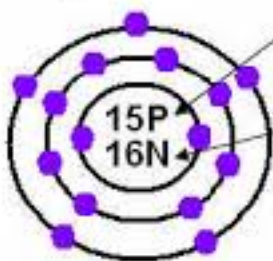
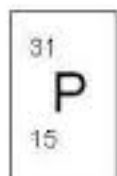


X = Mass number

Z = Atomic number

A = Element symbol

Example: phosphorus



# of Protons = atomic #

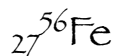
# of Protons = 15

# of Neutrons = mass # - atomic #

# of Neutrons = 31 - 15

# of Neutrons = 16

How did you figure it out?



Protons \_\_\_\_\_

\_\_\_\_\_

Neutrons \_\_\_\_\_

\_\_\_\_\_

Electrons \_\_\_\_\_

\_\_\_\_\_

Complete the following table.

Hint: Use a periodic table to find the name and symbols of the element!

Name of Element	Symbol	Atomic Number	Mass Number	# Protons	# Electrons	# Neutrons
oxygen						
	Hg					
		2				
			122			
				19		
					54	
		83				
	Cr					

Draw the Standard Notation for:

Nitrogen

Kr

Cu

Try p. 190 #1-3 in your textbook