STATION #1:
HOW BALANCED ARE YOU?

BALANCE THE FOLLOWING

EQUATIONS!

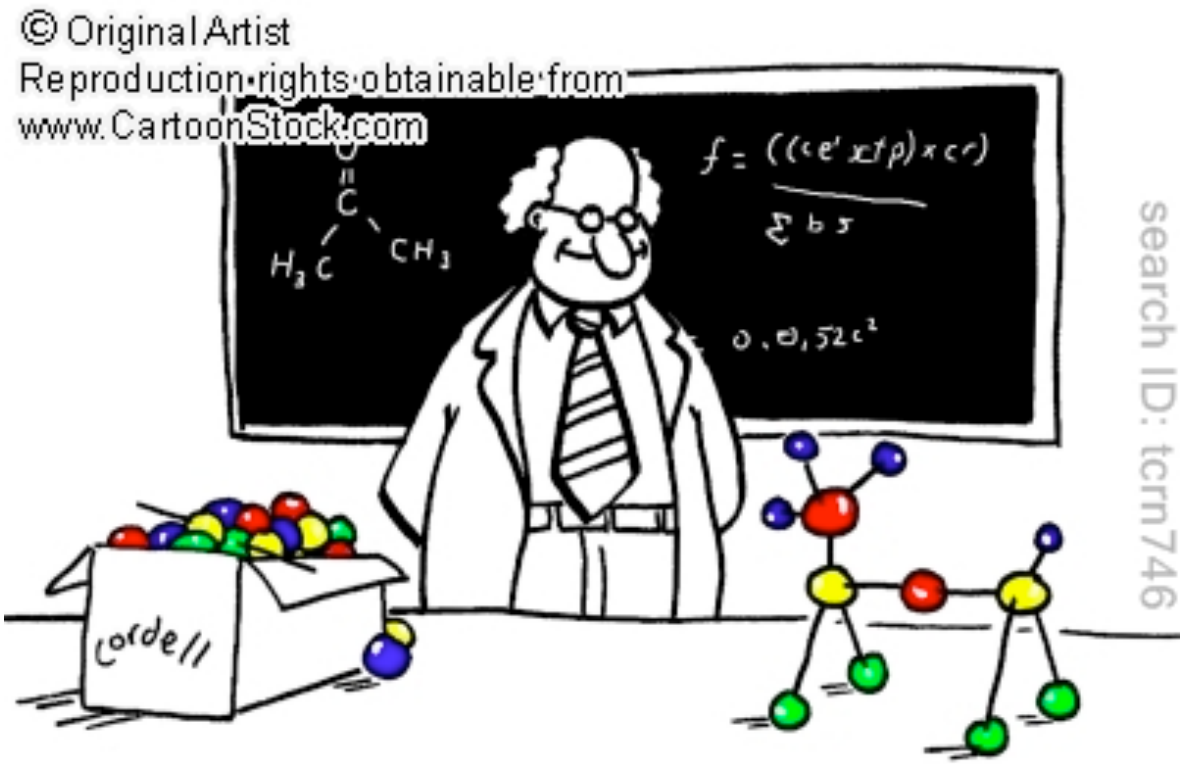
STATION #2:
WHAT'S IN A NAME?

WRITE THE IUPAC NAME FOR THE FOLLOWING CHEMICAL FORMULAS!

HINT: ARE THEY

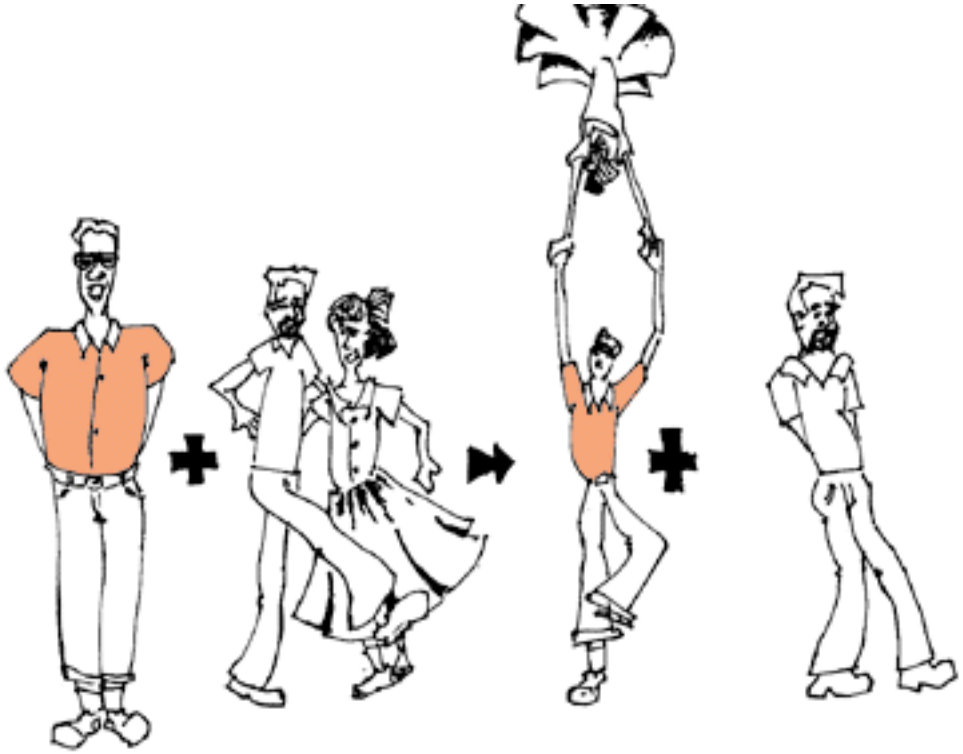
COVALENT OR IONIC?

STATION #3 : FORMULA FORRAY!



WRITE THE CHEMICAL FORMULAS FOR THE FOLLOWING COMPOUNDS USING IUPAC SYSTEM!

STATION #4 :
TO REACT OR NOT TO REACT...



WRITE THE TYPE OF CHEMICAL REACTION. CHOOSE FROM: SYNTHESIS, DECOMPOSITION, SINGLE/DOUBLE DISPLACEMENT, OR COMBUSTION.

STATION #5 : WORD, BRO!

WATCH YOUR NOMENCLATURE!

TURN THE FOLLOWING WORD EQUATIONS INTO CHEMICAL EQUATIONS.

STATION #1:
HOW BALANCED ARE YOU?

1) ___ C_3H_8 + ___ O_2 → ___ CO_2 + ___ H_2O

2) ___ Al + ___ Fe_3N_2 → ___ AlN + ___ Fe

3) ___ Na + ___ Cl_2 → ___ NaCl

4) ___ H_2O_2 → ___ H_2O + ___ O_2

5) ___ $C_6H_{12}O_6$ + ___ O_2 → ___ H_2O + ___ CO_2

6) ___ H_2O + ___ CO_2 → ___ C_7H_8 + ___ O_2

7) ___ $NaClO_3$ → ___ NaCl + ___ O_2

8) ___ $(NH_4)_3PO_4$ + ___ $Pb(NO_3)_4$ → ___$Pb_3(PO_4)_4$ + ___ NH_4NO_3

9) ___ BF_3 + ___ Li_2SO_3 → ___ $B_2(SO_3)_3$ + ___ LiF

10) ___ C_7H_{17} + ___ O_2 → ___ CO_2 + ___ H_2O

STATION #2:
WHAT'S IN A NAME?

1) $LiOH$ _____

2) PBr_3 _____

3) Na_2SO_4 _____

4) $(NH_4)_2S$ _____

5) $CaCO_3$ _____

6) CF_4 _____

7) $NaNO_3$ _____

8) P_2S_3 _____

9) $Al(NO_3)_3$ _____

10) $Mg(OH)_2$ _____

STATION #3 :
FORMULA FORRAY!

1) potassium oxide _____

2) phosphorus tribromide _____

3) calcium hydroxide _____

4) dinitrogen sulfide _____

5) carbon monoxide _____

6) diboron tetrahydride _____

7) phosphorus pentabromide _____

8) sulfur dichloride _____

9) sodium carbonate _____

10) aluminum acetate _____

STATION #4 :
TO REACT OR NOT TO REACT...

1) ____ NaBr + ____ $Ca(OH)_2$ → ___ $CaBr_2$ + ____ NaOH

Type of reaction: _____

2) ____ NH_3+ ____ H_2SO_4 → ____ $(NH_4)_2SO_4$

Type of reaction: _____

3) ____ C_5H_9O + ____ O_2 → ____ CO_2 + ____ H_2O

Type of reaction: _____

4) ____ Pb + ____ H_3PO_4 → ____ H_2 + ____ $Pb_3(PO_4)_2$

Type of reaction: _____

5) ____ HBr + ___ $Al(OH)_3$ → ___ H_2O + ___ $AlBr_3$

Type of reaction: _____

STATION #5 :
WORD, BRO!

1) Zinc and lead (II) nitrate react to form zinc nitrate and lead.

_____

2) Aluminum bromide and chlorine gas react to form aluminum chloride and bromine gas.

_____

3) Sodium phosphate and calcium chloride react to form calcium phosphate and sodium chloride.

_____

4) Potassium metal and chlorine gas combine to form potassium chloride.

_____

5) Aluminum and hydrochloric acid react to form aluminum chloride and hydrogen gas.